

Problematic text in the manuscript	Source found in the internet	Suggested action
alloys, steels and brass (Certified Reference Materials) A 0.1g amount of an alloy or steel or brass sample containing different composition of metals was accurately weighed and placed in a 50-mL Erlenmeyer flask. To it, 10-mL of concentrated HNO₃ and 1-mL of concentrated H₂SO₄ were carefully added and then covered with a watch-glass until the brisk reaction subsides. The solution was heated and simmered gently after the addition of another 5-mL of concentrated HNO₃ until all carbides were decomposed. The solution was carefully evaporated to dense white fumes to drive off the oxides of nitrogen and then cooled to room temperature (25±5)°C. After suitable dilution with deionized water, the contents of the Erlenmeyer flask were warmed to dissolve the soluble salts. The solution was then cooled and neutralized with a dilute NH₄OH solution in the presence of 1-2-mL of 0.01% (w/v) tartrate solution. The resulting solution filtered, if necessary, through Whatman no. 40 filter paper into a 25-mL calibrated flask. The residue (silica and tungstic acid) was washed with a small volume (5-mL) of hot (1:99) sulfuric acid, followed by water, the filtration and washing were collected in the same calibrated flask and the volume was made up to the mark with deionized water. A suitable aliquot (1-2-mL) of the above solution was taken into a 10-mL calibrated flask and the silver content was determined as described under general procedure using citrate or fluoride as masking agent. Based on five replicate analyses, the average silver concentration determined by spectrophotometric method was in good agreement with the spiked values.	http://ceacsu.edu.pk/PDF%20file/Vol%2013%20No%2001/22-35-PJAE-31012012-11.pdf	Kindly take the reference of the source inside your paper to avoid plagiarism complain.
Determination of silver in biological samples Regarding human blood (3-5 -mL) and urine (20- 50 -mL) were collected in polyethane bottles from the affected persons. The samples were taken into a 100-mL micro-Kjeldahl flask. A glass bead and 10 mL of concentrated nitric acid were added and the flask was placed on the digester under gentle heating following a method recommended by	http://pmcc.web-t.cisti.nrc.ca/articlerender.cgi?accid=PMC2701165	Kindly take the reference of the source inside your paper to avoid plagiarism

Stahr. ²⁹ When the initial brisk reaction was over, the solution was removed and cooled. 5 mL of concentrated HNO ₃ was added carefully, followed by the addition of 0.5 mL of 70 % HClO ₄ , and heating was continued to dense white fumes, repeating HNO ₃ addition if necessary. Heating was continued for at least 30 min; followed by cooling.		complain.
Determination of silver in soil samples An air-dried homogenized soil sample (100g) was accurately weighed and placed in a 100-mL micro-Kjeldahl flask. The sample was digested in the presence of an oxidizing agent following a method recommended by Jackson ³¹ . The content of flask was filtrated through Whatman No. 40 filter paper into a 25-mL calibrated flask, and neutralized with dilute ammonia in the presence of 1-2-mL of a 0.01% (w/v) EDTA solution. It was then diluted up to the mark with deionized water. Suitable aliquots (1-2-mL) were transferred into a 10-mL calibrated flask and the silver content was determined as described under general procedure using fluoride or thiocyanide solution as a masking agent and quantified from a calibration graph prepared concurrently.	http://ceacsu.edu.pk/PDF%20file/Vol ume%2013%20No%2001/22-35-PJAE-31012012-11.pdf	Kindly take the reference of the source inside your paper to avoid plagiarism complain.
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